

SECOND SEMESTER B.Sc. DEGREE EXAMINATION, MAY 2018

(CUCBCSS-UG)

Chemistry

CHE 2B 02 – THEORETICAL AND INORGANIC CHEMISTRY-II

Time : Three Hours

Maximum : 80 Marks

Section A (One word)*Answer all questions.**Each question carries 1 mark.*

1. The radial probability is _____ at the nucleus.
2. The Laplacian operator, $\nabla^2 =$ _____.
3. The second ionization energy is _____ than the first ionization energy.
4. Electron gain enthalpies of noble gases are _____.
5. The Born-Landé equation is _____.
6. The geometry of NH_3 is _____.
7. The ionic crystal will dissolve only when the solvation energy overcomes the _____.
8. If all the electrons in a molecule are paired, the substance is _____.
9. The higher the bond order, _____ is the bond.
10. In aqueous solution, HF dissociates to form _____ ion.

(10 × 1 = 10 marks)

Section B (Short Answer)*Answer any ten questions.**Each question carries 2 marks.*

11. What is meant by well-behaved wave function?
12. The atomic orbital 2s exists, but 2d does not. Explain why?
13. What is meant by eigen function and eigen value of an operator?
14. Electron gain enthalpy of Fluorine is lower than that of Chlorine. Justify.
15. The size of Cl^- is greater than that of Cl , while that of Na^+ is less than that of Na atom. Give reason.

Turn over

16. State Fajans rules.
17. What are the factors favouring the formation of ionic bond?
18. In PCl_5 axial bonds are longer than that of equatorial bonds. Explain.
19. The dipole moment of CO_2 is zero, while that of water is 1.84 D. Why?
20. Compare the bond energy of NO molecule, NO^+ ion and NO^- ion.
21. Why ortho nitrophenol is more volatile than para nitrophenol?
22. Differentiate Bonding and Anti bonding molecular orbitals.

(10 × 2 = 20 marks)

Section C (Paragraph)

Answer any **five** questions.

Each question carries 6 marks.

23. Briefly describe the postulates of quantum mechanics.
24. Write a note on quantum numbers and their significance.
25. Sketch the radial probability distribution curves for 2s, 3p and 3d orbitals. Explain.
26. Briefly describe Slater's rule and its applications.
27. Compare the properties of ionic and covalent compounds?
28. Discuss the hybridization and structure of XeF_6 .
29. Draw and explain molecular orbital diagram of B_2 molecule.
30. Write briefly on different types of induction and dispersion forces.

(5 × 6 = 30 marks)

Section D (Essays)

Answer any **two** questions.

Each question carries 10 marks.

31. Set up and solve the Schrödinger wave equation for a particle in a three dimension box and get expression for the wave function and energy.
32. (i) Describe the different electronegativity scales.
(ii) How does ionization energy of elements vary along a period and down a Group. Give reason.
33. Write a note on Born-Haber cycle and its application.
34. Describe the theories of metallic bond. How is the metallic properties explained, based on the theories?

(2 × 10 = 20 marks)